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Chapter 14 - Acids and Bases

**Chapter 14 (Acids and Bases)**

~~- Part 1 Acids and Bases~~

~~Chemistry - Basic~~

~~Introduction Acids Bases and~~

~~Salts Chapter 14 (Acids and Bases) - Part 4 Chapter 14~~

~~(Acids and Bases) - Part 2~~

~~Chapter 14 (Acids and Bases)~~

~~- Part 5  $K_a$   $K_b$   $K_w$  pH pOH pKa~~

~~pKb  $H^+$   $OH^-$  Calculations -~~

~~Acids \u0026 Bases, Buffer~~

~~Solutions , Chemistry Review~~

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Chapter 26 Fluid,  
Electrolyte, Acid-Base  
Balance

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Acid-Base Equilibria~~

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pH, pOH,  $H_3O^+$ ,  $OH^-$ ,  $K_w$ ,  $K_a$ ,  
 $K_b$ , pKa, and pKb Basic

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The Nature of Acids and  
Bases. A. Arrhenius Model 1.

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Acids produce hydrogen ions in aqueous solutions 2.

Bases produce hydroxide ions in aqueous solutions B.

Bronsted-Lowry Model 1.

Acids are proton donors 2.

Bases are proton acceptors

3. Chapter 14 - Acids and Bases -

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much of the acid base table  
as you can, as you read .  
Acid base conductivity and  
reactivity Conductivity,  
Reac'tivity,,  
Hydrochloric\*acid\* high high  
Acidic\*acid\* \*\* low\* medium

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Introduction. Originally the terms acid and base referred to taste. The practice of classifying substances according to their acidic (sour) or basic (alkaline or bitter) properties dates back to ancient times. An acid was something with a sour taste, such as lemon juice, and a base was something with a bitter taste, such as tonic water.

*Introduction to Acids and  
Bases (Worksheet) -  
Chemistry ...*

Acid and Base Worksheet -  
Answers. 1) Using your  
knowledge of the Brønsted-  
Lowry theory of acids and  
bases, write equations for

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the following acid-base reactions and indicate each conjugate acid-base pair: a)  $\text{HNO}_3 + \text{OH}^- \rightarrow \text{H}_2\text{O} + \text{NO}_3^-$ .  $\text{HNO}_3$  and  $\text{NO}_3^-$  make one pair  $\text{OH}^-$  and  $\text{H}_2\text{O}$  make the other. b)  $\text{CH}_3\text{NH}_2 + \text{H}_2\text{O} \rightarrow \text{CH}_3\text{NH}_3^+ + \text{OH}^-$

*Acid and Base Worksheet -  
Answers - Chemistry Made  
Easy*

Chapter 14 Acids And Bases  
an acid that is composed of  
H, O, and a third element.  
The H & O are covalently  
bonded and are called the  
hydroxyl group. Property of  
bases. taste bitter, feel  
slippery, change color of  
indicators, react with acids  
to produce a salt and water

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neutralization rxn, are electrolytes. Arrhenius or traditional base.

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CHAPTER 14 REVIEW . Acids  
and Bases. SHORT ANSWER

Answer the following  
questions in the space  
provided. 1. a. Write the  
two equations that show the  
two-stage ionization of  
sulfurous acid in water.

stage 1:  $\text{H}_2\text{SO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l})$   
 $\sim \text{H}_3\text{O}^+(\text{aq}) + \text{HSO}_3^-(\text{aq})$  b.

Which stage of ionization  
usually produces more ions?  
Explain your answer.

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