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Changes:

Exothermic or
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Bomb Calorimeter
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17

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As liquids absorb heat at their boiling points, the temperature remains constant

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mochemistry
while they
vaporize. true

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CHAPTER 17,

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(continued) Use

the heating

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mechemistry
shown below to
answer Questions
5, 6, and 7.

**section 17.1 the
flow of energy
heat and work**

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2. 0.100 g of H₂
and an excess of
O₂ are

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compressed into
a bomb
calorimeter
containing 1200
g of water. The
temperature
before the
reaction is
25.00 °C, and
after the
reaction it goes
to 27.16 °C.

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Thermochemistry **Packet**

CHAPTER 17,
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(continued) Use
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curve for water
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5, 6, and 7.

Heating Curve
for Water

Boiling point

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Melting E point

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Heat supplied 5.

Label the

melting point

and boiling

point

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the graph. 6.

What happens to

the temperature

during melting

and

vaporization?

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10. Complete the
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combustion of
natural gas. Use
the

thermochemical
equation in the
first paragraph
on page 517 as a
guide. SECTION
17.3 HEAT IN
CHANGES OF STATE
(pages 520–526)

This section
explains heat
transfers that

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mechanistry
occur during
melting,
freezing,
boiling, and
condensing.

SECTION 17.1 THE FLOW OF ENERGY HEAT AND WORK (pages 505–510)

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Kiddy Math

1. How much energy must be absorbed by 20.0 g of water to increase its temperature from 283.0 °C to 303.0 °C? 2.

When 15.0 g of steam drops in temperature from 275.0 °C to 250.0 °C, how

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much heat energy
is released?
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